

Double Replacement Reaction Worksheet

What happens when solutions in each row is mixed with solutions in each column?

- If a reaction occurs and a precipitation is observed, write *precipitate forms*.
- If the mixed solutions remain clear, write *no reaction*.

	$K_2CO_3(aq)$	$NaCl(aq)$	$Na_3PO_4(aq)$
$Mg(NO_3)_2(aq)$			
$Ca(NO_3)_2(aq)$			
$Sr(NO_3)_2(aq)$			
$Ba(NO_3)_2(aq)$			

For each reaction that occurred, write:

1. the *balanced* chemical reactions, and
2. the net ionic reactions and identify the spectator ion(s).

Remember to include physical states.