

Tutorial 0010

1. There are two stable isotopes of bromine. Isotope A has a mass of 78.918336 amu and isotope B has a mass of 80.916289 amu.
 - Determine the natural abundance of the isotopes, and report the answer to three significant figures.

 - Write the symbol for the two isotopes.
2. There are two naturally occurring isotopes of silver, ^{107}Ag and ^{109}Ag . The atomic mass of silver is 107.87 amu, with 51.82% of the atoms being ^{107}Ag . The mass of an atom of ^{107}Ag is 106.9 amu. What is the mass, in amu, of an atom of ^{109}Ag .
3. Write the electron configuration of the ground state of calcium.
4. What is the maximum number of the electrons possible in the third energy levels of atoms?
5. List three orbitals and describe their shapes.
6. Naming Compounds. See problems at the end of UNIT 6.
<http://nobel.scas.bcit.ca/chem0010/unit6/pnomena.htm>
<http://nobel.scas.bcit.ca/chem0010/unit6/pnomenb.htm>

