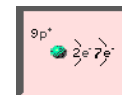


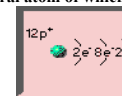
## Assignment 10

Choose/Fill-in the correct answer.

- An atom has an atomic number of 29 and a mass number of 64. The number of electrons in the atom is
  - 29
  - 64
  - 35
  - 93
- Lewis electron-dot symbols consist of symbol representing the element and the arrangement of dots which represent
  - the atomic number
  - the core electrons
  - the valence electrons
  - the ground state of the element
- Which of the following electron configuration is the ground state of manganese?
  - $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^5$
  - $1s^2 2s^2 2p^6 3s^2 3p^6 3d^7$
  - $[Ar]3d^7$
  - $1s^2 2s^2 2p^6 3s^2$
- The number of electrons contained in  $^{56}\text{Fe}^{2+}$  ion is
  - 56
  - 26
  - 28
  - 24
- The maximum number of electrons possible in the fourth energy levels of atoms is
  - 8
  - 18
  - 32
  - 64
- The maximum number of electrons possible in the fourth energy levels of atoms is
  - 8
  - 18
  - 32
  - 64
- If an element has the following electronic arrangement of



it is

- an alkali metal.
  - a halogen.
  - a noble gas.
  - an alkaline-earth metal.
- Calculate the number of protons and neutrons in  $^{36}\text{Ar}$ .
    - protons = 18; neutrons = 16
    - protons = 18; neutrons = 18
    - protons = 18; neutrons = 22
    - protons = 18; neutrons = 36
  - The Bohr diagram shown below is a neutral atom of which element?
 
    - sodium
    - magnesium
    - aluminum
    - silicon
  - Orbitals describe the probability regions for finding the electrons. The orbital that matches the dumbbell shape is:
    - s
    - p
    - d
    - s and d

Send to obtain your score